

CANDIDATE
NAME

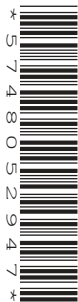
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CENTRE
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CHEMISTRY

5070/22

Paper 2 Theory

October/November 2015

1 hour 30 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Section A

Answer **all** questions.

Write your answers in the spaces provided in the Question Paper.

Section B

Answer any **three** questions.

Write your answers in the spaces provided in the Question Paper.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

A copy of the Periodic Table is printed on page 20.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **20** printed pages.

Section A

Answer **all** the questions in this section in the spaces provided.

The total mark for this section is 45.

A1 Choose from the following elements to answer the questions below.

aluminium
argon
carbon
copper
iodine
iron
lead
magnesium
nitrogen
oxygen
sulfur

Each of these elements can be used once, more than once or not at all.

Which element

(a) has an ion which, in aqueous solution, reacts with aqueous sodium hydroxide to give a red-brown precipitate,

.....[1]

(b) has an atom with an electronic configuration with only five occupied electron shells,

.....[1]

(c) has an oxide which decolourises acidified potassium manganate(VII),

.....[1]

(d) has a sulfate which is insoluble in water,

.....[1]

(e) provides an inert atmosphere for the extraction of reactive metals,

.....[1]

(f) produces ammonia when it is warmed with an aqueous mixture of sodium nitrate and sodium hydroxide?

.....[1]

[Total: 6]

A2 Plants contain many different types of carbon compounds.

(a) Carbon compounds are made in plants by photosynthesis.

Write the overall equation for photosynthesis and state the conditions required for this process.

.....
.....
.....[3]

(b) Starch is a polymer found in plants.

Starch can be hydrolysed.

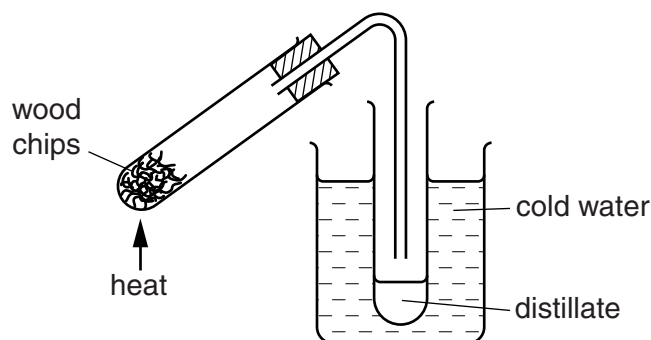
(i) Give the name of the product formed when starch is hydrolysed.

.....[1]

(ii) Give the reagent and conditions needed to hydrolyse starch.

.....[2]

(c) When wood chips are heated in the absence of air, a colourless distillate is formed.



The distillate contains ethanoic acid, esters and other organic compounds.

(i) Ethanoic acid reacts with calcium hydroxide to form a salt and water.

Give the name and formula of the salt formed.

name

formula [2]

(ii) One of the esters in the distillate has the molecular formula, $C_4H_8O_2$.

Draw the structure of an ester with this formula, showing all the atoms and all the bonds.

[1]

(iii) The distillate also contains a compound with the following composition.

element	percentage by mass
carbon	37.5
hydrogen	12.5
oxygen	50.0

Deduce the empirical formula of this compound.

empirical formula[2]

[Total: 11]

A3 Water for use in the home is treated using carbon and chlorine.

(a) Explain the purpose of using carbon and chlorine in water treatment.

carbon.....

chlorine [2]

(b) In some parts of the world, drinking water is purified by desalination.

What is meant by the term *desalination*?

..... [1]

(c) River water may contain pollutants from agricultural sources.
These pollutants may cause eutrophication.

Give the names of two anions present in fertilisers which contribute to eutrophication.

..... and [1]

(d) An aqueous solution of barium chloride is added to a sample of water which contains sulfate ions. A white precipitate forms.

Construct an ionic equation, including state symbols, for this reaction.

..... [2]

[Total: 6]

A4 Iron is extracted in a blast furnace. The raw materials required are

- iron ore, which contains iron(III) oxide, Fe_2O_3 ,
- limestone,
- coke (carbon),
- air.

- (a)** The coke first burns in air to form carbon dioxide.
The carbon dioxide is then reduced by coke to produce carbon monoxide.
The carbon monoxide reduces the iron(III) oxide to iron.

Write equations for

- (i)** the reduction of carbon dioxide by coke,

.....[1]

- (ii)** the reduction of iron(III) oxide to iron by carbon monoxide.

.....[1]

- (b)** Why is limestone added to the blast furnace?

.....[1]

- (c)** Another ore of iron contains an oxide with the formula Fe_3O_4 .

Calculate the percentage by mass of iron in Fe_3O_4 .

[2]

- (d) Iron can be obtained by the electrolysis of an aqueous acidified solution of iron(II) sulfate. The reactions at the electrodes are given.



Which reaction is oxidation and which is reduction? Explain your answer.

.....
.....
.....[2]

- (e) Attaching magnesium strips to the iron hulls of ships can lower their rate of rusting.

Explain how the magnesium stops the iron from rusting.

.....
.....
.....
.....[2]

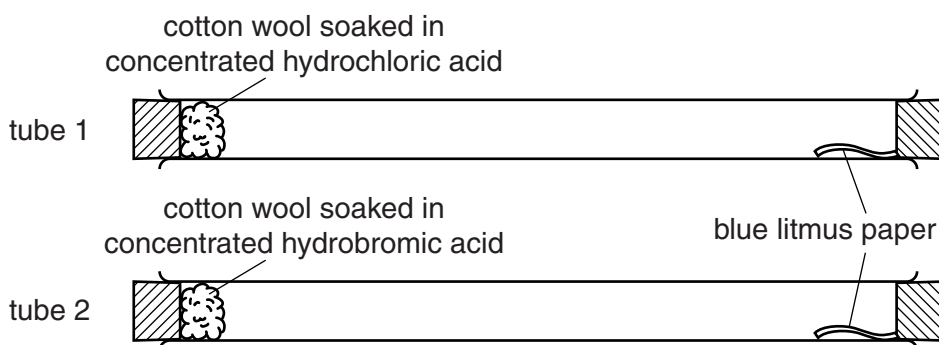
- (f) Aqueous iron(II) chloride is one of the products formed when iron reacts with hydrochloric acid.

Construct an equation for this reaction.

.....[1]

[Total: 10]

A5 (a) Two students set up tubes as shown.



Concentrated hydrochloric acid produces fumes of hydrogen chloride.
 Concentrated hydrobromic acid produces fumes of hydrogen bromide.

Four minutes after setting up the experiment, the litmus paper in tube 1 turns red.
 Seven minutes after setting up the experiment, the litmus paper in tube 2 turns red.

Use the kinetic particle theory to explain

(i) how the gases move through the tubes,

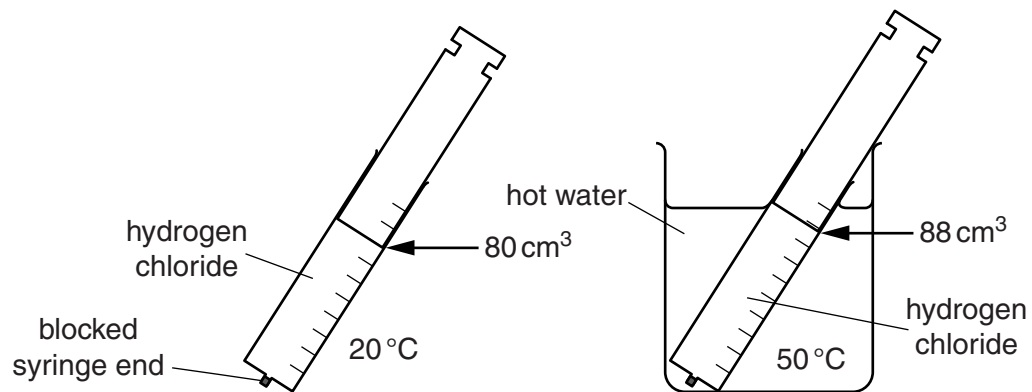
.....

 [2]

(ii) why the gases take different times to reach the litmus paper.

.....
 [1]

- (b) A gas syringe is filled with 80 cm^3 of hydrogen chloride gas at 20°C . The syringe is placed in some hot water at 50°C . The atmospheric pressure does not change but the volume of the gas in the syringe increases to 88 cm^3 .



Use the kinetic particle theory to explain why the volume increases.

.....

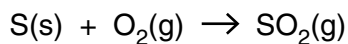
.....

.....

.....[2]

[Total: 5]

A6 When one mole of sulfur burns, 247 kJ of energy is released.



(a) Calculate the energy released when 9.60 g of sulfur is burnt.

[2]

(b) When sulfur dioxide is passed through aqueous sodium hydroxide, sodium hydrogensulfite is formed. Sodium hydrogensulfite contains the hydrogensulfite ion, HSO_3^- .

Construct the ionic equation for this reaction.

.....[1]

(c) The hydrogensulfite ion is a weaker acid than ethanoic acid.

Samples of 0.1 g of magnesium are added separately to 0.1 mol/dm^3 ethanoic acid and 0.1 mol/dm^3 hydrogensulfite ions.

Explain why the reaction is faster with ethanoic acid.

.....

[2]

(d) Calculate the volume of 0.10 mol/dm^3 sodium hydroxide which contains 3.2 g of sodium hydroxide.

volume of aqueous sodium hydroxide = cm^3 [2]

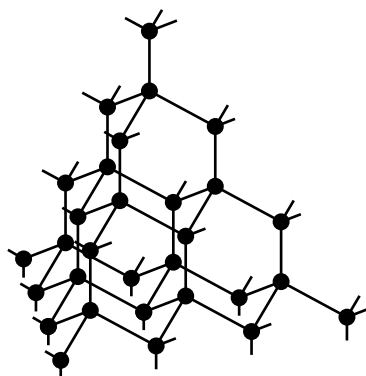
[Total: 7]

Section B

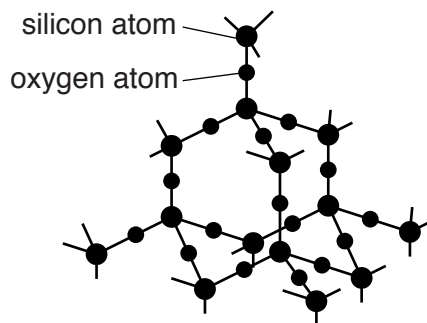
Answer **three** questions from this section in the spaces provided.

The total mark for this section is 30.

B7 The structures of diamond and silicon dioxide are shown.



diamond



silicon dioxide

(a) Describe one similarity in the structures of diamond and silicon dioxide.

.....[1]

(b) Explain why silicon dioxide has a high melting point.

.....

.....[2]

(c) Silicon dioxide reacts with hot concentrated sodium hydroxide to form sodium silicate, Na_2SiO_3 .

Deduce the formula of the silicate ion.

.....[1]

(d) Explain why diamond does not conduct electricity.

.....[1]

(e) Graphite is a form of carbon which conducts electricity.
Graphite is used as electrodes in electrolysis.

What is the meaning of the term *electrolysis*?

.....

.....[1]

- (f) Aluminium is extracted by the electrolysis of molten aluminium oxide dissolved in cryolite.

Construct the equation for the reaction at

- (i) the anode (positive electrode),

.....[1]

- (ii) the cathode (negative electrode).

.....[1]

- (g) An isotope of silicon is represented by the symbol ${}_{14}^{29}\text{Si}$.

Deduce the number of protons and neutrons in this isotope.

number of protons

number of neutrons[1]

- (h) Silicon has a relative atomic mass of 28.

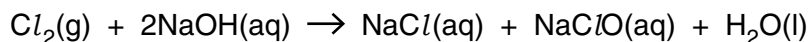
Define the term *relative atomic mass*.

.....

.....[1]

[Total: 10]

B8 Chlorine reacts with sodium hydroxide.



- (a) A volume of 144 cm³ of chlorine gas, measured at room temperature and pressure, is passed into 38.0 cm³ of 0.250 mol/dm³ sodium hydroxide.

Show by calculation which reactant is in excess.

[3]

- (b) Chlorine displaces bromine from an aqueous solution of potassium bromide.

- (i) Construct the equation for this reaction.

.....[2]

- (ii) Explain why bromine does not react with aqueous potassium chloride.

.....
[1]

- (iii) Give the charge of a chloride ion and its electronic configuration.

charge

electronic configuration[2]

- (c) Explain why sodium chloride does not conduct electricity when solid but does conduct electricity when molten.

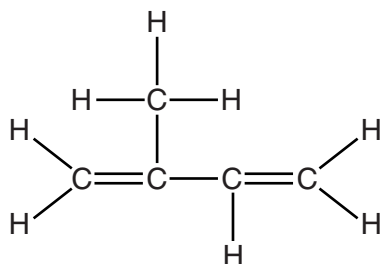
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[2]

[Total: 10]

Question B9 begins on page 16.

- B9** When rubber is heated in the absence of air, a small amount of isoprene is formed. The structure of isoprene is shown.



- (a)** Isoprene is an unsaturated hydrocarbon.

Describe a test for an unsaturated hydrocarbon.

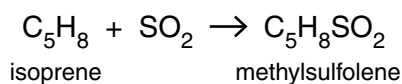
.....
 [2]

- (b)** Isoprene is a liquid at 25°C.

Describe the arrangement and motion of the particles in isoprene at 25°C.

.....
 [2]

- (c)** Isoprene reacts with sulfur dioxide to form methylsulfolene.



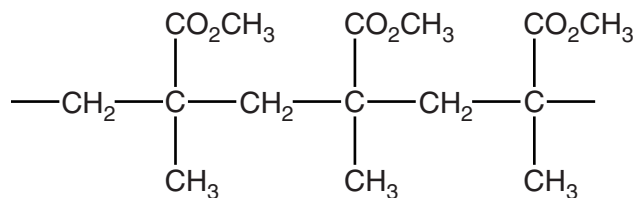
Calculate the maximum mass of methylsulfolene that can be formed from 100 g of isoprene.

mass of methylsulfolene = g [3]

- (d) (i) What feature of the isoprene molecule is responsible for it forming an addition polymer?

.....[1]

- (ii) *Perspex* is also an addition polymer.
The diagram shows part of the polymer chain of *Perspex*.

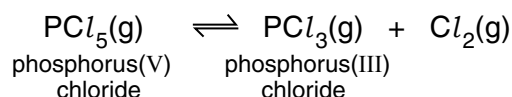


Deduce the formula of the monomer used to make *Perspex*.

[2]

[Total: 10]

B10 At 200°C and 200 atmospheres pressure, phosphorus(V) chloride forms an equilibrium mixture with phosphorus(III) chloride and chlorine.



- (a) Predict and explain the effect of decreasing the pressure on the position of this equilibrium. The temperature remains constant.

.....

 [2]

- (b) Predict and explain the effect of increasing the concentration of chlorine on the position of this equilibrium.

.....

 [2]

- (c) The table shows the percentage of phosphorus(III) chloride in the equilibrium mixture at different temperatures. The pressure is the same in each case.

temperature/°C	% PCl ₃ in the mixture
200	48
300	95
400	99

- (i) Describe how the composition of this equilibrium mixture changes with temperature.

.....
 [1]

- (ii) Explain what this tells you about the energy change in this reaction.

.....
 [1]

(d) How is the position of equilibrium affected by the presence of a catalyst?

.....[1]

(e) The rate of this reaction increases with increase in temperature.
Explain why.

.....
.....
.....
.....[2]

(f) Phosphorus(V) chloride reacts with water. Phosphoric acid, H_3PO_4 , and hydrogen chloride are formed.

Construct the equation for this reaction.

.....[1]

[Total: 10]

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DATA SHEET The Periodic Table of the Elements

Group																																																																																																													
I	II	III	IV	V	VI	VII	O																																																																																																						
		1 H Hydrogen 1					4 He Helium 2																																																																																																						
7 Li Lithium 3	9 Be Beryllium 4		11 B Boron 5	12 C Carbon 6	13 Al Aluminium 13	14 N Nitrogen 7	15 O Oxygen 8	16 F Fluorine 9	17 Ne Neon 10	19 Ar Argon 18	20 K Potassium 19	21 Ca Calcium 20	22 Sc Scandium 21	23 Ti Titanium 22	24 V Vanadium 23	25 Cr Chromium 24	26 Mn Manganese 25	27 Fe Iron 26	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36	37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54	55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	58 Ce Cerium 58	59 Pr Praseodymium 59	60 Nd Neodymium 60	61 Pm Promethium 61	62 Sm Samarium 62	63 Eu Europium 63	64 Gd Gadolinium 64	65 Tb Terbium 65	66 Dy Dysprosium 66	67 Ho Holmium 67	68 Er Erbium 68	69 Tm Thulium 69	70 Yb Ytterbium 70	71 Lu Lutetium 71	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86	87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89	90 Th Thorium 90	91 Pa Protactinium 91	92 U Uranium 92	93 Np Neptunium 93	94 Pu Plutonium 94	95 Am Americium 95	96 Cm Curium 96	97 Bk Berkelium 97	98 Cf Californium 98	99 Es Einsteinium 99	100 Fm Fermium 100	101 Md Mendelevium 101	102 No Nobelium 102	103 Lr Lawrencium 103	104 Rf Rutherfordium 104	105 Db Dubnium 105	106 Sg Seaborgium 106	107 Bh Bohrium 107	108 Hs Hassium 108	109 Mt Meitnerium 109	110 Ds Darmstadtium 110	111 Rg Roentgenium 111	112 Cn Copernicium 112	113 Nh Nihonium 113	114 Fl Flerovium 114	115 Mc Moscovium 115	116 Lv Livermorium 116	117 Ts Tennessine 117	118 Og Oganesson 118

* 58–71 Lanthanoid series
† 90–103 Actinoid series

a	a = relative atomic mass
X	X = atomic symbol
b	b = atomic (proton) number

Key

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).